

Ionic Bonding Review

Name _____

Period _____ Date _____

Part I: Completion

negative ions	positive polyatomic	metal cubic gas	electron configuration anion	electron cation
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1. An ionic bond is an attraction between oppositely charged _____.
2. Anions have a _____ charge.
3. An atom becomes an ion by losing or gaining _____.
4. The _____ rule states that atoms tend to gain, lose or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of argon _____.
6. The atoms that make up _____ ions are bonded together by ionic bonds.

Part II, True or False - If the statement is true, write "true." If it is false, change the underlined word or words to make it true.

7. Calcium becomes a monatomic cation by gaining two electrons.
8. A cation has a positive charge.
9. A binary ionic compound contains only one kind of cation and one kind of anion.
10. Monatomic ions consist of more than one atom.
11. The Lewis dot diagram for chlorine, a group 17 (7A) element, has six electrons.
12. Ionic compounds have very low melting points.

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Answers To Chemthink Ionic Formulas

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Answers To Chemthink Ionic Formulas

In this Chemthink tutorial, you will explore ionic formulas and take a short quiz. Topics include: positive and negative ions; balancing positive and negative charges; writing ionic formulas; The following fill-in-the-blank notes were provided by Rebecca Carlock at Centennial High School. Thank you Rebecca! Fill In The Blank Notes PDF Version MS Word Version: Thank you so much to Mr. Charles ...

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Chemthink Ionic Bonding Worksheet Answer Key

In this Chemthink tutorial, you will explore ionic bonding and take a short quiz. Topics include: how ions form and why they stick together; the 3-dimensional shape of an ionic crystal; counting ions in a crystal to determine the ionic formula; This new HTML5 version was built from the ground up to run on iPads, mobile phones, Chromebooks, and real computers in any HTML5-capable browser ...

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