

26. Which of the following electron configurations is correct?
- A) Ga: $[\text{Kr}]3d^9 4s^2 4p^2$
 B) Me: $(\text{Kr})5s^2 4d^1$
 C) Ca: $[\text{Ar}]4s^2 3d^9$
 D) Br: $[\text{Kr}]3d^9 4s^2 4p^7$
 E) Rf: $[\text{Xe}]6s^2 4f^{14} 5d^9 6p^7$
27. The electron configuration of Tl^{3+} is
- A) $[\text{Ar}]4s^2$
 B) $[\text{Ar}]4s^2 3d^1$
 C) $[\text{Ar}]3d^1$
 D) $[\text{Ar}]4s^2 3d^2$
 E) none of these
28. $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^2$ is the correct electron configuration for which of the following atoms?
- A) Ca
 B) Ti
 C) Ge
 D) Zr
 E) none of these
29. Which of the following have 10 electrons in the d orbitals?
- A) Mn
 B) Fe
 C) Cu
 D) Zn
 E) two of the above
30. Order the elements S, Cl, and F in terms of increasing atomic radii.
- A) S, Cl, F
 B) Cl, F, S
 C) F, S, Cl
 D) F, Cl, S
 E) S, F, Cl
31. Which of the following is the highest energy orbital for a silicon atom?
- A) 1s
 B) 2s
 C) 3s
 D) 3p
 E) 3d
32. Copper exhibits the expected electron configuration.
- A) True
 B) False
33. List the following atoms in order of increasing ionization energy: Li, Na, C, O, F.
- A) $\text{Li} < \text{Na} < \text{C} < \text{O} < \text{F}$
 B) $\text{Na} < \text{Li} < \text{C} < \text{O} < \text{F}$
 C) $\text{F} < \text{O} < \text{C} < \text{Li} < \text{Na}$
 D) $\text{Na} < \text{Li} < \text{F} < \text{O} < \text{C}$
 E) $\text{Na} < \text{Li} < \text{C} < \text{F} < \text{O}$

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Practice Exercise 2 (11.1) In which of the following substances is significant hydrogen bonding possible: methylene chloride (CH₂Cl₂), phosphine (PH₃), hydrogen peroxide (HOOH), or acetone (CH₃COCH₃)? AP Chemistry Chapter 11. Intermolecular Forces, Liquids, and Solids - 7 - Comparing Intermolecular Forces • Dispersion forces - found in all substances, strength dep. on molecular ...

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