

AP CHEMISTRY

MRS. SPENCER

CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY1) Which of the following are weak electrolytes? HCl, HC₂H₃O₂, NH₃, KCl

A) HCl, KCl

D) HC₂H₃O₂, NH₃B) HCl, HC₂H₃O₂, NH₃, KClE) HCl, HC₂H₃O₂, KClC) HC₂H₃O₂, KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and HNO₃(aq)?A) K⁺ and H⁺ C) K⁺ and NO₃⁻ E) OH⁻ onlyB) H⁺ and OH⁻ D) H⁺ and NO₃⁻

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of CaCO₃ when aqueous solutions ofNa₂CO₃ and CaCl₂ are mixed is _____.A) 2Na⁺(aq) + CO₃²⁻(aq) → Na₂CO₃(aq)

E)

B) 2Na⁺(aq) + 2Cl⁻(aq) → 2NaCl(aq)Na₂CO₃(aq) + CaCl₂(aq) → 2NaCl(aq) + CaCO₃(s)C) Na⁺(aq) + Cl⁻(aq) → NaCl(aq)D) Ca²⁺(aq) + CO₃²⁻(aq) → CaCO₃(s)

Answer: D Diff: 3

4) Which of the following are strong acids? HI, HNO₃, HF, HBr

A) HF, HBr

C) HI, HF, HBr

E) HI, HNO₃, HBrB) HI, HNO₃, HF, HBrD) HNO₃, HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces _____.

A) water and a salt

C) oxygen gas

E) ammonia

B) hydrogen gas

D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na

B) Au

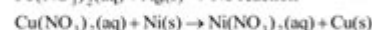
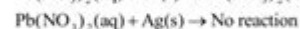
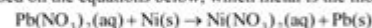
C) Fe

D) Ca

E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni

B) Ag

C) Cu

D) Pb

E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100

B) 0.0500

C) 0.0333

D) 0.0250

E) 125

Answer: A Diff: 3

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AP Chemistry - Santa Ana Unified School District

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