

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO ₃	Chloric acid	H ₃ O ⁺	ClO ₃ ⁻¹	Chlorate
H ₂ SO ₃	Sulfurous acid	2H ₃ O ⁺	SO ₃ ⁻²	Sulfite
H ₂ SO ₄	Sulfuric acid	2H ₃ O ⁺	SO ₄ ⁻²	Sulfate
H ₂ PO ₃	Phosphorous acid	3H ₃ O ⁺	PO ₃ ⁻³	Phosphite
H ₃ PO ₄	Phosphoric acid	3H ₃ O ⁺	PO ₄ ⁻³	Phosphate
HNO ₃	Nitric acid	H ₃ O ⁺	NO ₃ ⁻¹	Nitrate
HNO ₂	Nitrous acid	H ₃ O ⁺	NO ₂ ⁻¹	Nitrite
H ₂ CO ₃	Carbonic acid	2H ₃ O ⁺	CO ₃ ⁻²	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO₃? *ClO₃⁻¹ has the name of chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not HClO₃.*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.
- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



2
Key

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AP WORKSHEET 06d: Acids & Bases I STRONG acids and bases. (Assume $pOH + pH = 14$). 1. Calculate the $[OH^-]$ for a 0.0545 M solution of Hydrochloric Acid (HCl). (2) 2. Calculate the pOH for a 0.0333 M solution of lithium hydroxide. (2) 3. Calculate the $[H^+]$ for a 0.0745 M solution of calcium hydroxide assuming complete dissociation of the compound. (2) 4. Complete the following table. (6 ...

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Acid Base Titration Simulation: Answers: 08CDEFGHIJ (8.3 - 8.5, 8.7 - 8.10) Weak Acids, Titrations and Buffers: Answers: 08ABCDEF GHIJ (8.1 - 8.10) Acid & Bases Summary: Answers: 08ABCDEF GHIJAP (8.1 - 8.10) FRQ's 08ABCDEF GHIJ (LO MAPPED) Answers: 08ABCDEF GHIJAPU (8.1 - 8.10) FRQ's 07ABCDEF GHIJ KLMN (UNMAPPED) Answers

AP WORKSHEETS 9 UNITS - Adrian Dingle's Chemistry Pages

The Chemistry of Acids & Bases 2 In each of the acid examples---notice the formation of H_3O^{+} --- this species is named the hydronium ion. It lets you know that the solution is acidic! (hydronium, H_3O^{+} riding piggy-back on a water molecule; water is polar and the + charge of the "naked" proton is greatly attracted to Mickey's chin!)

AP*Chemistry The Chemistry of Acids and Bases

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Mrs. Rick's Website - Worksheets

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Acid and Base Worksheet - Studylib

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Baker, Mrs. (Science) / AP Chemistry

AP FREE RESPONSE QUESTIONS - ACIDS/BASES 1994 D A chemical reaction occurs when 100. milliliters of 0.200-molar HCl is added dropwise to 100. milliliters of 0.100-molar Na₃PO₄ solution. (a) Write the two net ionic equations for the formation of the major products. (b) Identify the species that acts as both a Brønsted acid and as a Brønsted base in the equation in (a). Draw the Lewis ...

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