

Answer key

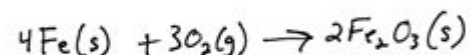
LOHS Honors Chemistry  
Balancing Chemical Equations

Part I: Balance these chemical equations. Record what type of reaction on the blank in the right margin.

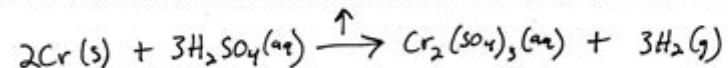
- |  |           |
|--|-----------|
| 1. $\text{BaCl}_2 + 2\text{KClO}_3 \rightarrow \text{Ba}(\text{ClO}_3)_2 + 2\text{KCl}$                        | <u>DR</u> |
| 2. $\text{Pb}(\text{OH})_2 \rightarrow \text{PbO} + \text{H}_2\text{O}$  | <u>D</u>  |
| 3. $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$                                    | <u>SR</u> |
| 4. $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + 2\text{KNO}_3$              | <u>DR</u> |
| 5. $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 2\text{Al}(\text{OH})_3 + 3\text{CaSO}_4$ | <u>DR</u> |
| 6. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$  | <u>SR</u> |
| 7. $\text{Sn} + 2\text{KOH} \rightarrow \text{K}_2\text{SnO}_2 + \text{H}_2$                                   | <u>SR</u> |
| 8. $2\text{PaI}_5 \rightarrow 2\text{Pa} + 5\text{I}_2$  | <u>D</u>  |
| 9. $3\text{SiF}_4 + 3\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SiF}_6 + \text{H}_2\text{SiO}_3$          | <u>SR</u> |
| 10. $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_4$                                    | <u>S</u>  |
| 11. $16\text{Na} + \text{S}_8 \rightarrow 8\text{Na}_2\text{S}$  | <u>S</u>  |
| 12. $\text{N}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_2$                                      | <u>S</u>  |

Part II: Write balanced chemical equations for each of the reactions below. Assume that all ionic compounds are dissolved in water (aqueous). Also be sure to include physical states where appropriate.

1. Iron metal combines with oxygen gas to form iron (III) oxide solid



2. Chromium metal and sulfuric acid ( $\text{H}_2\text{SO}_4$ ) react to form chromium (III) sulfate solution and hydrogen gas.



3. Barium chromate solution and aqueous hydrochloric acid (HCl) react to produce barium chloride solution and hydrogen chromate (solid).



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# Balancing Chemical Equations Gap Fill Exercise Answers

## Balancing Chemical Equations Gap Fill

### Balancing Chemical Equations Gap Fill

Balancing Chemical Equations Gap-fill exercise. Fill in all the gaps, then press "Check" to check your answers. Use the "Hint" button to get a free letter if an answer is giving you trouble. You can also click on the "[?]" button to get a clue. Note that you will lose points if you ask for hints or clues! 1.  $H_2 + O_2 \rightarrow H_2O$  2.  $H_2 + N_2 \rightarrow NH_3$  3.  $Al_2O_3 \rightarrow Al + O_2$  4.  $AlCl_3 \rightarrow Al$  ...

### Balancing Chemical Equations - ScienceGeek.net

Balancing Chemical Equations Gap-fill exercise. Fill in all the gaps, then press "Check" to check your answers. Use the "Hint" button to get a free letter if an answer is giving you trouble. You can also click on the "[?]" button to get a clue. Note that you will lose points if you ask for hints or clues! 1.  $H_2 + O_2 \rightarrow H_2O$  2.  $H_2 + N_2 \rightarrow NH_3$  3.  $Al_2O_3 \rightarrow Al + O_2$  4.  $KClO_3 \rightarrow KCl$  ...

### Balancing Chemical Equations - ScienceGeek.net

Unit 4 Benchmark #2 - Balancing Equations (Version B) Gap-fill exercise. Enter your answers in the gaps. Every space will require a coefficient. Unlike when we balance equations in class, you will have to include coefficients of "one" by typing in a value of "1." When you have entered all the answers, click on the "Check" button. 1.  $CaO + H_2O \rightarrow Ca(OH)_2$  2.  $Ca + O_2 \rightarrow CaO$  3.  $Cu + S \rightarrow CuS$  ...

### Unit 4 Benchmark #2 - Balancing Equations (Version B)

State symbols like (aq), (g), (l) and (s) have been left out so you can concentrate on deducing the balancing numbers from 2, 3, 4 etc. Some are very easy but some are much more tricky and the last one is ghastly! (1) .  $HCl + CuO \rightarrow CuCl_2 + H_2O$

### Gap-fill worksheet question Quiz "Balancing Chemical ...

balancing chemical equations gap fill exercise fill in all the gaps then press check to check your answers use the hint button to get a free letter if an answer is giving you trouble you can also click on the button to get a clue unit 4 benchmark 2 balancing equations version b gap fill exercise enter your answers in the gaps every space will require a coefficient unlike when we balance ...

### Balancing Chemical Equations Gapfill Exercise Answers

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### Balancing Chemical Equations Worksheet Gap Fill Exercise ...

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### Balancing Chemical Equations Gap Fill Exercise Answers ...

To balance a chemical equation, enter an equation of a chemical reaction and press the

Balance button. The balanced equation will appear above. Use uppercase for the first character in the element and lowercase for the second character. Examples: Fe, Au, Co, Br, C, O, N, F. Ionic charges are not yet supported and will be ignored.

### Chemical Equation Balancer

Balancing Chemical Equations

### Balancing Chemical Equations

Instructions on balancing chemical equations: Enter an equation of a chemical reaction and click 'Balance'. The answer will appear below; Always use the upper case for the first character in the element name and the lower case for the second character.

Examples: Fe, Au, Co, Br, C, O, N, F. Compare: Co - cobalt and CO - carbon monoxide;

To enter an electron into a chemical equation use {-} or e

### Balance Chemical Equation - Online Balancer

We found some Images about Balancing Chemical Equations Worksheet Gap Fill Exercise:

### Balancing Chemical Equations Worksheet Gap Fill Exercise ...

A chemistry tutorial designed to help learn the basic principles of balancing chemical equations, along with examples and methods of balancing different chem...

### Balancing Chemical Equations - Chemistry Tutorial - YouTube

The balanced chemical equation has 16 carbon atoms, 36 hydrogen atoms, and 50 oxygen atoms on each side. Balancing equations requires some practice on your part as well as some common sense. If you find yourself using very large coefficients or if you have spent several minutes without success, go back and make sure that you have written the formulas of the reactants and products correctly.

### 7.4: How to Write Balanced Chemical Equations - Chemistry ...

### Chemical Equations Home Learning Worksheet GCSE | Teaching ...

Balancing chemical equation with substitution. Practice: Balancing chemical equations 1. This is the currently selected item. Next lesson. Stoichiometry. Balancing chemical equation with substitution. Our mission is to provide a free, world-class education to anyone, anywhere. Khan Academy is a 501(c)(3) nonprofit organization. Donate or volunteer today! Site Navigation. About. News; Impact ...

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### Balancing Chemical Equations Worksheet Answer Key ...

Balancing Chemical Equations Gapfill exercise Enter your answers in the gaps. Every space will require a coefficient. Unlike when we balance equations in class, you will have to include coefficients of "one" by typing in a value of "1." When you have entered all the answers, click on the "Check" button. 1.

### Balancing Chemical Equations | Pearltrees

Erika Tan balances three sample chemical equations for all levels of chemistry.

TRANSCRIPT: So I have 3 chemical equations here that I need to balance today. The

first one is blank  $N_2 + blank H_2 \dots$

### How to Balance Chemical Equations

UNIT 8 - CHEMICAL EQUATIONS 3 Worksheets - Honors 10. zinc + nickel (II) nitrate 11. magnesium + silver sulfate 12. potassium sulfate + aluminum chlorate potassium chlorate + aluminum sulfate 13. ethane (C 2 H 6) + oxygen carbon dioxide + water 14. copper (II) chloride + sodium phosphate

### BALANCING EQUATIONS PRACTICE WORKSHEET

If the equation is not balanced, change the coefficients of the molecules until the number of atoms of each element on either side of the equation balance. Check that the atoms are in fact balanced. (we will look at this a little later): Add any extra details to the equation e.g. phase symbols. Worked example 1: Balancing chemical equations 1

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