

Name _____ Date _____ Class _____

14 THE BEHAVIOR OF GASES

Chapter Test A

A. Matching

Match each description in Column B with the correct term in Column A. Write the letter of the correct description in the blank.

- | Column A | Column B |
|--|--|
| _____ 1. ideal gas pressure (P) | a. The volume of a fixed mass of gas is directly proportional to the Kelvin temperature if the volume is kept constant. |
| _____ 2. Boyle's law | b. As constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases. |
| _____ 3. Dalton's law of partial pressures | c. $n = \frac{PV}{RT}$ |
| _____ 4. ideal gas law | d. An ideal gas is a theoretical model for an ideal gas. |
| _____ 5. reduced gas law | e. A gas tends to react in a ratio of their stoichiometric coefficients with the reactants in a balanced equation. |
| _____ 6. Charles' law | f. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$ |
| _____ 7. Avogadro | g. $P = T + a + b + c$ |
| _____ 8. partial pressure | h. For a given mass of gas at constant temperature, the volume of gas varies inversely with the pressure. |

B. Multiple Choice

Circle the best answer and write the letter in the box.

- _____ 9. As the temperature of a fixed volume of gas increases, the pressure will
 a. not change. b. increase.
 c. decrease. d. remain the same.
- _____ 10. A breathing mixture used by divers is three moles helium, oxygen, and carbon dioxide. What is the partial pressure of oxygen in (a) 200 kPa total pressure and $V_{O_2} = 6.0 \text{ L}$ and $V_{CO_2} = 12.0 \text{ L}$?
 a. 30 kPa b. 20 kPa
 c. 17 kPa d. 26.7 kPa

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Under what conditions of temperature and pressure is the behavior of real gases most like that of ideal gases? a. low temperature and low pressure c. high temperature and low pressure b. low temperature and high pressure d. high temperature and high pressure ___ 50. If the atmospheric pressure on Mt. Everest is one-third the atmospheric pressure at sea level, the partial pressure of oxygen on ...

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Real gases deviate from ideal behavior at low temperatures and high pressures. What is an ideal gas? An ideal gas is a gas that follows the gas laws at all conditions of pressure and temperature. Determine the volume occupied by 0.582 mol of a gas at 15°C if the pressure is 81.8 kPa. 17.0 L. What pressure is exerted by 0.450 mol of a gas at 25°C if the gas is in a 0.650-L container? 1.71 ...

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