

Chapter 10 States of Matter

Chapter Opener

__Chapter Overview, TE Review the objectives listed in the Student Edition.

Section 1 The Kinetic-Molecular Theory of Matter

PACING: 45 minutes

PENNSYLVANIA ACADEMIC STANDARDS FOR SCIENCE AND TECHNOLOGY:

3.1.B.1 Distinguish between different types of models and modeling techniques and apply their appropriate use in specific applications;

3.1.C.3 Examine and describe physical patterns in motion;

3.4.A.4 Describe phases of matter according to the Kinetic Molecular Theory;

3.4.B.4 Use knowledge of conservation of energy and momentum to explain common phenomena.

Objectives

1. **State** the kinetic-molecular theory of matter, and describe how it explains certain properties of matter.
2. **List** the five assumptions of the kinetic-molecular theory of gases. Define the terms ideal gas and real gas.
3. **Describe** each of the following characteristic properties of gases: expansion, density, fluidity, compressibility, diffusion, and effusion.
4. **Describe** the conditions under which a real gas deviates from "ideal" behavior.

FOCUS (5 minutes)

__ **Lesson Starter, TE** Open a bottle of perfume or ammonia at the front of the classroom and ask students to raise their hands when they first detect the odor. Ask them to explain the event in terms of molecules.

MOTIVATE (10 minutes)

__ **Discussion, TE** To begin the discussion of the differences between ideal and real gases, have students discuss daily activities in which real and ideal situations exist.

TEACH (20 minutes)

__ **PowerPoint**

__ **Visual Strategy, Figure 1, TE** Have students relate the assumptions of the kinetic-molecular theory of gases to this figure.

__ **Visual Strategy, Figure 3, TE** Ask students to infer what will happen if the pressure on the gas continues to increase.

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Chapter 13 States Of Matter

For this chapter, you will need to know about kinetic energy. Kinetic energy is what is created when an object moves. Kinetic theory also states that all matter consists of particles that are constantly in motion. Therefore, when in a closed container, all these moving particles that make up, well, everything will have to hit each other. That ...

Chapter 13: States of Matter - Chemistry by Anna

A theory that explains the states of matter, based on the concept that all matter consists of tiny particles that are in constant motion

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Chapter 13 States of Matter 137 SECTION 13.1 THE NATURE OF GASES (pages 385-389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the

Name Date Class STATES OF MATTER 13

Figure 13.3: The images are trying to convey that a solid (table salt), liquid (water), and gas (helium) all have the same _____ kinetic energy at the same temperature (20 C), even though the three substances are in different _____.

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km gases 7.notebook 1 February 05, 2015 Feb 109:15 AM Chapter 13 "States of Matter" Feb 109:15 AM •OBJECTIVES: •Describe the assumptions of the "kinetic theory" as it

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Chapter 13 States Of Matter Work Answer Key

-The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the particles in a gas are rapid, constant, and random.

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Next Answer Chapter 13 - States of Matter - 13.4 Changes of State - Chemistry & You - Page 439: Q Previous Answer Chapter 13 - States of Matter - 13.3 The Nature of Solids - 13.3 Lesson Check - Page 434: 23. Answers by Chapter; Chapter 1 - Introduction to Chemistry; Chapter 2 - Matter and Change; Chapter 3 - Scientific Measurement ;

Chapter 4 - Atomic Structure; Chapter 5 - Electrons in Atoms ...

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Chapter 13 States Of Matter the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you. For example, you can easily recognize the shape of your desk; you know that your backpack cannot hold seven textbooks. Chapter 13: States of Matter Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 ...

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Chapter 13 "States of Matter" Section 13.1 The Nature of Gases. OBJECTIVES: Describe the assumptions of the "kinetic theory" as it applies to gases. Section 13.1 The Nature of Gases. OBJECTIVES: Interpret gas pressure in terms of kinetic theory. Section 13.1 The Nature of Gases. OBJECTIVES: Define the relationship between Kelvin temperature and average kinetic energy. Section 13.1 The ...

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In this chapter, your exploration of the states of matter will go far beyond everyday, casual observations. You will explore a characteristic shared by liquids and gases, and investigate how these substances produce pressure. You'll meet the physics principles that explain how huge wooden ships can float on water, and how enormous metal aircraft can fly. You will find out why some solids ...

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Intermolecular forces, including dispersion forces, dipole-dipole forces, and hydrogen bonds, determine a substance's state at a given temperature. The particles in solids and liquids have a limited range of motion and are not easily compressed. Matter changes phase when energy is added or removed. Chapter 13: gases

Chapters 12 & 13: States of Matter and Gases - ANNE ...

Chapter 13 States Of Matter Worksheet Title: Chapter 13 States of Matter 1 Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Chapter 13 States Of Matter Worksheet Look at the text on page 315 for the answer. You are already familiar with the three ...

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