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12

## THE BEHAVIOR OF GASES

## Conceptual Curriculum

- Concrete concepts
- More abstract concepts or math/problem-solving

## Standard Curriculum

- Core content
- Extension topics

## Honors Curriculum

- Core honors content
- Options to accelerate

SECTION 12.1 THE PROPERTIES OF GASES (pages 327–328)   

This section describes the properties of gas particles and explains how the kinetic energy of gas particles relates to Kelvin temperature.

## ► Kinetic Theory Revisited (pages 327–328)

1. What theory explains the behavior of gases with respect to conditions such as temperature and pressure? kinetic theory
2. If you notice that a sealed bag of potato chips bulges when placed near a sunny window, what can you hypothesize about the relationship between the temperature and pressure of an enclosed gas?  
An increase in temperature causes an increase in the pressure exerted by the gas.
3. List three basic assumptions of the kinetic theory about the properties of gases.
  - a. Gas particles are so small in relation to the distances between them that their individual volumes can be assumed to be insignificant.
  - b. No attractive or repulsive forces exist between gas particles.
  - c. Gas particles are in constant random motion, traveling in independent straight paths until they collide with each other or other objects.
4. Circle the letter next to each sentence that is true concerning the compressibility of gases.
  - a. The large relative distances between gas particles means that there is considerable empty space between them.
  - b. The assumption that gas particles are far apart explains gas compressibility.
  - c. Compressibility is a measure of how much the volume of matter decreases under pressure.
  - d. Energy is released by a gas when it is compressed.
5. Look at Figure 12.1 on page 327. How does an automobile air bag protect the crash dummy from being broken as a result of impact?  
The gases used to inflate the airbag are compressible and are able to absorb a considerable amount of energy.

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# Chapter 14 Properties Of Gases Answers

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### **Chapter 14 1 The Properties Of Gases Answers**

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File Type PDF Chapter 14 The Properties Of Gases Answers SECTION 14.1 PROPERTIES OF GASES(pages 413-417) 14.02 If any transaction for the purposes of this Chapter is also a connected transaction for the purposes of Chapter 14A, the listed issuer will, in addition to complying with the provisions of this Chapter, have to comply

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how gas pressure is affected by the amount of gas, its volume, and its temperature. Chapter 14 1 The Properties Of Gases Answers Chapter 14 1 The Properties Of Gases Answers HC Verma Solutions Vol 1 Some Mechanical Properties of Matter Chapter 14. Question 1: A load of 10 kg is suspended by a metal wire 3 m long and having a cross-sectional area 4 mm<sup>2</sup>. Find (a) the stress (b) the strain

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### **Chapter 14 Properties Of Gases Answers**

SECTION 14.1 PROPERTIES OF GASES(pages 413-417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature. Compressibility (pages 413-414) 1. Look at Figure 14.1 on page 413. SECTION 14.1 PROPERTIES OF GASES(pages 413-417)

**14 The Behavior Of Gases Chapter Quiz - aplikasidapodik.com**  
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### **Chapter 14 (& 13.1) - Gases - KEIO ACADEMY OF NEW YORK ...**

Let us start by reviewing some properties of gases. Gases have no definite shape or volume; they tend to fill whatever container they are in. They can compress and expand, sometimes to a great extent. Gases have extremely low densities, one-thousandth or less the density of a liquid or solid. Combinations of gases tend to mix together spontaneously; that is, they form solutions. Air, for example, is a solution of mostly nitrogen and oxygen. Any understanding of the properties of gases must ...

### **Introduction to Gases - Introductory Chemistry - 1st ...**

Properties of Gases - easily compressible into smaller volumes - increase pressure= increased density - exert pressure on their surroundings - pressure is required to contain a gas-expand without limits - completely and uniformly occupy the volume of a container-Diffuse into one another - gases in container mix completely - do not separate upon standing-Described in the following terms ...

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