Name _____ Date _____ Class _____

CHAPTER 16 STUDY GUIDE

Reaction Rates

Section 16.1 A Model for Reaction Rates

In your textbook, read about expressing reaction rates and explaining reactions and their rates.

Use each of the terms below just once to complete the passage.

collision theory	activated complex	mol/(L·s)
activation energy	reaction rate	

According to the **(1)**_____, atoms, ions, and molecules must collide in order to react. Once formed, the **(2)**_____ is a temporary, unstable arrangement of atoms that may then form products or may break apart to reform the reactants. Every chemical reaction requires energy, and the minimum amount of energy that reacting particles must have to form the activated complex is the **(3)**_____. In a chemical reaction, the **(4)**_____ is the change in concentration of a reactant or product per unit time. It may be expressed using the units of **(5)**_____.

Use the energy diagram for the rearrangement reaction of methyl isonitrile to acetonitrile to answer the following questions.

6. What kind of reaction is represented by this diagram, endothermic or exothermic?

_____

7. What is the chemical structure identified at the top of the curve on the diagram?

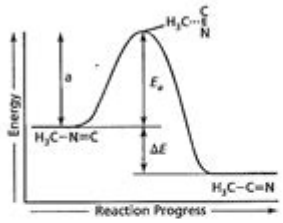
_____

8. What does the symbol E_a represent?

_____

9. What does the symbol ΔE represent?

_____



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Reaction Rates Chapter 16 • Reaction Rates 559 SStart-Up Activitiestart-Up Activities LAUNNCH CH LLabab How can you accelerate a reaction? Some chemical reactions go so slowly that nothing seems to be happening. In this lab, you can investigate one way of speeding up a slow reaction. Procedure 1. Read and complete the lab safety form. 2. Create a Before and After table to record your ...

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Writing Reaction Rate Laws Calculating a Reaction Rate Example Reaction Order Example #1 Rate Law: show the relationship between the rate of a chemical reaction and the concentration of reactants at a given temperature The following reaction is first order in hydrogen and second

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