

Chapter 17 Thermochemistry Study Guide

17.1 – 17.2 Thermochemical Equations

Human Chemistry

- Make the following conversions:
 - 444 cal to joules = 1.86×10^3 J
 - 1.8 kJ to joules = 1.8×10^3 J
 - 0.45 kJ to calories = 1.1×10^2 cal
- Classify each of these processes as endothermic or exothermic:
 - condensing steam – **exo**
 - burning alcohol – **exo**
 - evaporating alcohol – **endo**
 - baking a potato – **endo**
- The specific heat capacity for silver is 0.24 J/g°C. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(107.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

- It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C. Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

- A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C. The equilibrium temperature in the calorimeter is 21.8°C. Calculate the specific heat capacity of copper; assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

- A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C. Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is 0.444 J/g°C.

$$\begin{aligned} -15.0(.444)(T_f - 100.0) &= 55.0(4.184)(T_f - 23.0) \\ -6.66T_f + 666 &= 230.12T_f - 5292.76 \\ 5958.76 &= 236.78 T_f \\ 25.2^\circ\text{C} &= T_f \end{aligned}$$

- Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at 95°C and then adds 15.8 L of water at 75°C, what will the final temperature of the water be?

$$\begin{aligned} -20(T_f - 95) &= 15.8(T_f - 75) \\ -20T_f + 1900 &= 15.8 T_f - 1185 \\ 3085 &= 35.8 T_f \end{aligned}$$

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CHAPTER 17, Thermochemistry (continued) GUIDED PRACTICE PROBLEM 12 (page 513) 12. When 50.0 mL of water containing 0.50 mol HCl at 22.5°C is mixed with 50.0 mL of water containing 0.50 mol NaOH at 22.5°C in a calorimeter, the temperature of the solution increased to 26.0°C. Guided Reading Key Ch17 - MAFIADOC.COM

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Acces PDF Chapter 17 Thermochemistry Practice Problems Answers Thermochemistry Answers Key Thermochemistry Practice Problems (Ch. 6) 1. Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20 °C. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach 21 °C? Explain your

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Chapter 17 Thermochemistry 437 Practice Problems In your notebook, solve the following problems. SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK Use the three-step problem-solving approach you learned in Chapter 1. 1. How many kilojoules of energy are in a donut that contains 200.0 Calories? 2.

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CH 8: Specific Heat Problems Worksheet . 1. How much energy must be absorbed by 20.0 g of water to increase its temperature from 283.0 °C to 303.0 °C? 2. When 15.0 g of steam drops in temperature from 275.0 °C to 250.0 °C, how much heat energy is released? 3. How much energy is required to heat 120.0 g of water from 2.0 °C to 24.0 °C? 4. If 720.0 g of steam at 400.0 °C absorbs 800.0 kJ ...

Thermochemistry Problems - Worksheet Number One

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