

Name _____ Date _____ Class _____

16-2 Practice Problems (continued)

17. Write the equilibrium expression for the following reaction.

$$\text{H}_2\text{CO}_3(\text{s}) \rightleftharpoons \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$$
18. Write the equilibrium expression for the following reaction.

$$\text{CO}_2(\text{g}) + \text{H}_2(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{l})$$
-
19. At 740°C, $K_{\text{eq}} = 0.0060$ for the decomposition of calcium carbonate (CaCO_3), which is described by the equation

$$\text{CaCO}_3(\text{s}) \rightleftharpoons \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$$

 Find Q and predict how the reaction will proceed if $[\text{CO}_2] = 0.0004 \text{ M}$.
20. For the reaction

$$\text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{CO}_2(\text{g})$$

 $K_{\text{eq}} = 5.10$ at 527°C. If $[\text{CO}] = 0.15 \text{ M}$, $[\text{H}_2\text{O}] = 0.25 \text{ M}$, $[\text{H}_2] = 0.42 \text{ M}$, and $[\text{CO}_2] = 0.37 \text{ M}$, calculate Q and determine how the reaction will proceed.
21. At 340°C, $K_{\text{eq}} = 0.064$ for the reaction

$$\text{Fe}_2\text{O}_3(\text{s}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{Fe}(\text{s}) + 3\text{H}_2\text{O}(\text{g})$$

 Given that $[\text{H}_2] = 0.45 \text{ M}$ and $[\text{H}_2\text{O}] = 0.37 \text{ M}$, find Q and predict how the reaction will proceed.
22. At 2130°C, $K_{\text{eq}} = 0.0025$ for the reaction

$$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$$

 If $[\text{N}_2] = 0.81 \text{ M}$, $[\text{O}_2] = 0.75 \text{ M}$, and $[\text{NO}] = 0.030 \text{ M}$, find Q and determine the direction in which the reaction will proceed.
23. Ammonia is synthesized from nitrogen and hydrogen in the reaction

$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$$

 At 500°C, the equilibrium constant for this reaction is 0.080. Given that $[\text{NH}_3] = 0.0596 \text{ M}$, $[\text{N}_2] = 0.600 \text{ M}$, and $[\text{H}_2] = 0.420 \text{ M}$, find Q and predict how the reaction will proceed.
24. The decomposition of antimony pentachloride (SbCl_5) is described by the equation

$$\text{SbCl}_5(\text{g}) \rightleftharpoons \text{SbCl}_3(\text{g}) + \text{Cl}_2(\text{g})$$

 At 448°C, the equilibrium constant for this reaction is 0.0251. What is the value of Q if $[\text{SbCl}_5] = 0.095 \text{ M}$, $[\text{SbCl}_3] = 0.020 \text{ M}$, and $[\text{Cl}_2] = 0.050 \text{ M}$? How will this reaction proceed?
25. At 1000°C, $K_{\text{eq}} = 1.0 \times 10^{-13}$ for the decomposition of hydrofluoric acid (HF), as described in the reaction

$$2\text{HF}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{F}_2(\text{g})$$

 If $[\text{HF}] = 23.0 \text{ M}$, $[\text{H}_2] = 0.540 \text{ M}$, and $[\text{F}_2] = 0.380 \text{ M}$, determine the value of Q and predict how the reaction will proceed.
26. At 1227°C, K_{eq} for the following reaction is 0.15.

$$2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$$

 If $[\text{SO}_2] = 0.344 \text{ M}$, $[\text{O}_2] = 0.172 \text{ M}$, and $[\text{SO}_3] = 0.056 \text{ M}$, find Q and determine how the reaction will proceed.

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