

## CHAPTER 3 SOLUTIONS MANUAL

## Matter—Properties and Changes

## Section 3.1 Properties of Matter

pages 70–75

## Problem-Solving Lab

1. Explain why the flow of a compressed gas must be controlled for practical and safe use.

The flow of compressed gas must be controlled to control the amount and the rate at which gas is released.

2. Predict what would happen if the valve on a full tank of compressed gas were suddenly opened all the way or if the tank were accidentally punctured.

Without the regulator device, the gas would rush out of the tank with a force powerful enough to transform the tank into a dangerous, uncontrolled projectile.

## Section 3.1 Assessment

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1. Create a table that describes the three common states of matter in terms of their shape, volume, and compressibility.

	Volume	Shape	Compressibility
Solid	Definite	Definite	Incompressible
Liquid	Definite	Takes shape of container and fills container to the extent of its own volume	Virtually incompressible
Gas	Fills volume of container	Takes shape of container	Compressible

2. Describe the characteristics that identify a sample of matter as being a substance.

The sample of matter must have a uniform and unchanging composition to be a substance.

3. Classify each of the following as a physical or chemical property.
  - a. Iron and oxygen form rust.  
chemical

- b. Iron is more dense than aluminum.  
physical

- c. Magnesium burns brightly when ignited.  
chemical

- d. Oil and water do not mix.  
physical

- e. Mercury melts at  $-39^{\circ}\text{C}$ .  
physical

4. **Organize** Create a chart that compares physical and chemical properties. Give two examples for each type of property.

The chart should make clear that physical properties can be observed without changing the composition of the sample, which is not the case for chemical properties. Mass and density are examples of physical properties. Fermentation and rusting are examples of chemical properties.

## Section 3.2 Changes in Matter

pages 76–79

## Practice Problems

page 78

5. Use the data in the table to answer the following questions.

Aluminum and Liquid Bromine Reaction		
	Before Reaction	After Reaction
Aluminum	10.3 g	0.0 g
Liquid bromine	100.0 g	8.5 g
Compound	0.0 g	

How many grams of bromine reacted? How many grams of compound were formed?

$$\text{amount of bromine that reacted} = 100.0 \text{ g} - 8.5 \text{ g} = 91.5 \text{ g}$$

$$\text{amount of compound formed} = 100.0 \text{ g} + 10.3 \text{ g} - 8.5 \text{ g} = 101.8 \text{ g}$$

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# Chapter 3 Supplemental Problems Matter Properties And Changes Answer Key

## [Chapter 3 Supplemental Problems Matter](#)

### Chapter 3 Supplemental Problems Matter

Supplemental Problems Answer Key Chemistry: Matter and Change 49 5 3 100% 5 18% Mass reactants 5 Mass products Mass nitrogen 1 Mass hydrogen 5 Mass ammonia Mass nitrogen 5 Mass ammonia 2 Mass hydrogen 14 g 5 17 g 2 3 g Chapter 4 1. Use the periodic table to complete the following table. Atomic number 5 number of protons 5 number of electrons 2.

### Chapter 3 - HONORS CHEMISTRY

Chapter 3 Accelerated Motion 6 c. b. What is his acceleration between t 60.0 s and t 61.0 s? fi fi 2 0.0m/s 3.0m/s 61.0 s 60.3s 3.0m/s v a t vv tt d. Assuming constant acceleration, how far did he walk during the first 5 s? fi fi 2 ii 22 1.5m/s 0.0m/s 10.0 s 0.0 s 1 2 2 a 1 ( Supplemental Problems Teacher Support continued

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### Chapter 3 Supplemental Problems Matter Properties And ...

Chapter 3 SG 3.1 SG 3.2 SG 3.4 Chapter 3 Supplemental Problems Chapter 3 Review Physical and Chemical Changes Lab Chapter 5 SG 5.1 Flame Test Lab SG 5.2 The Wave and Particle Nature of Light Bohr's Model/Quantum Mechanical Model Orbital Diagrams SG 5.3 Abbreviated Configurations Using Electron Configurations Chapter 5

Supplemental Problems ...

### Chapter 3 Supplemental Problems Answer Key Physics

Supplemental Problems for Chapter 3 1) Consider a structure with space group P4 2 cm (no. 101), and  $c = 4a$ . (i) Name the point group from which this space group is derived, and sketch a diagram illustrating the general equivalent positions. (ii) Explain the orientation and action of the operators 4 2, c, and m in this space group.

### Supplemental Problems for Chapter 3

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### CHAPTER 3 Matter—Properties and Changes

Supplemental Problems Chemistry: Matter and Change • Chapter 3 3 MatterMatter—Properties and ChangesProperties and Changes 1. An 18-g sample of element A combines com-pletely with a 4-g sample of element B to form the compound AB. What is the mass of the compound formed? 2. A substance breaks down into three component elements when it is heated. The mass of each

### Supplemental Problems - MARRIC

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Appendix A Supplemental Practice Problems. . . . . 871 ... 874 Chemistry: Matter and Change Practice Problems Chapter 6 Section 6-2 1. Identify the group, period, and block of an atom with the following electron configuration. a. [He]2s22p1 b. [Kr]5s24d5 c. [Xe]6s25f146d5 2. Write the electron configuration for the element fitting each of the following descriptions. a. The noble ...

### Appendix A: Supplemental Practice Problems

Chapter Assessment - Chapter 3 — Matter—Properties and Changes Reviewing Vocabulary 1. f 2. d 3. e 4. a 5. c 6. b 7. i 8. g 9. h 10. k 11. j 12. Both are characteristics of substances. A physical property can be observed without changing the composition of the substance. A chemical property is the ability or tendency of a substance to change

into another substance. 13. Both are kinds of ...

### Chapter Assessment - Chapter 3 — Matter—Properties and ...

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