

Name _____ Class _____ Date _____

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

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Chapter 6 Review Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

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Chapter 6 Review- Chemical Bonding Send article as PDF . A chemical bond between atoms results from the attraction between the valence electrons and ____ of different atoms. nuclei. A covalent bond consists of. a shared electron pair, if two covalently bonded atoms are identical, the bond is identified as. nonpolar covalent . A covalent bond is which there is an unequal attraction for the ...

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CHAPTER 6 REVIEW Chemical Bonding SECTION 6-2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule. 3.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. ____ The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. ____ In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions. 3. ...

CHAPTER 6 REVIEW Chemical Bonding

CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT ANSWER Answer the following questions in the space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn't water a linear molecule?

CHAPTER 6 REVIEW Chemical Bonding - clarkchargers.org

Review Previous Concepts Chemical Bonding CHAPTER 6 Section 1 Introduction to Chemical Bonding What is a chemical bond and why does it form? Section 2 Covalent Bonding and Molecular Compounds What is a molecular formula? What are the characteristics of a covalent bond? How does a Lewis structure represent a molecule? Section 3 Ionic Bonding and Ionic Compounds What is a crystal lattice? How ...

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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the

following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms.

Chapter 6 Review Chemical Bonding Mixed Answers

CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms.

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