

KEY

Chemistry 12

Unit 2 - Chemical Equilibrium

14. Explain why a syringe containing NO<sub>2</sub> gas will first get *darker* and *then lighter* in colour when compressed. Use the equilibrium equation:

① 
$$\text{N}_2\text{O}_4(\text{g}) + \text{heat} \rightleftharpoons 2\text{NO}_2(\text{g})$$

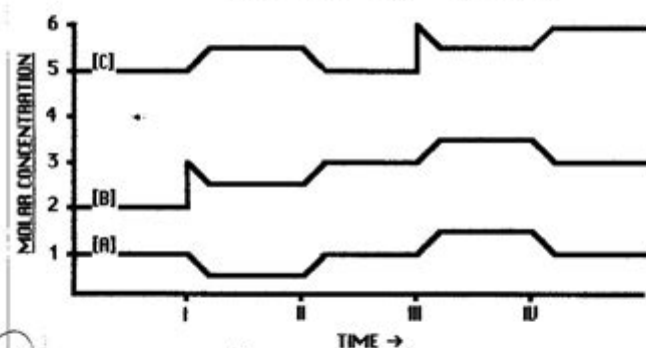
$$\text{colourless} \qquad \qquad \qquad \text{brown}$$
 darker first due to increased conc. of both gases including NO<sub>2</sub>  
 Equil<sup>m</sup> will shift left to counteract inc. pressure so [NO<sub>2</sub>] (brown) will decrease and the mixture will be lighter

15. Explain why a flask containing NO<sub>2</sub> will get *lighter* in colour when put into ice water. Use the equation:

① 
$$\text{N}_2\text{O}_4(\text{g}) + \text{heat} \rightleftharpoons 2\text{NO}_2(\text{g})$$

$$\text{colourless} \qquad \qquad \qquad \text{brown}$$
 Equil<sup>m</sup> will shift left to counteract dec. in temp, so [NO<sub>2</sub>] (brown) will decrease.

16. Given the following graph showing the concentrations of species A, B and C, state what changes in **temperature** or **concentration** are responsible for each of the shifts shown on the graph. The equilibrium equation is:



- a) At time I, the B is added
- b) At time II, the temp is increased
- c) At time III, the C is added
- d) At time IV, the temp is decreased

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