

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_



## Overview Chemical Bonds

**Directions:** All of the statements below are false as written. In the space provided, write a term or phrase that makes the statement true when it is substituted for the underlined words.

- \_\_\_\_\_ 1. The properties of a compound are the same as the properties of the elements that it contains.
- \_\_\_\_\_ 2. Superscript numbers in chemical formulas tell how many atoms of each element are found in a unit of compound.
- \_\_\_\_\_ 3. All the noble gases except helium have 18 electrons in their outer energy level.
- \_\_\_\_\_ 4. A(n) chemical formula is the force that holds atoms together in a compound.
- \_\_\_\_\_ 5. An ion is a(n) neutral particle that has either more or fewer electrons than protons.
- \_\_\_\_\_ 6. Oxidation numbers are written as subscripts.
- \_\_\_\_\_ 7. A(n) covalent bond is the force of attraction between the opposite charges of the ions in an ionic compound.
- \_\_\_\_\_ 8. The charge on a compound is always positive.
- \_\_\_\_\_ 9. Equal sharing of electrons in covalent bonds results in polar molecules.
- \_\_\_\_\_ 10. Only two identical atoms can share electrons unequally.
- \_\_\_\_\_ 11. A binary compound contains five different elements.
- \_\_\_\_\_ 12. An oxidation number tells how many protons an atom must gain, lose, or share to become stable.
- \_\_\_\_\_ 13. The oxidation number of the copper(II) ion is 3+.
- \_\_\_\_\_ 14. When writing chemical formulas, add superscripts so that the sum of the oxidation numbers equals ten.
- \_\_\_\_\_ 15. A polyatomic ion never has a positive or negative charge.
- \_\_\_\_\_ 16. The polyatomic ion  $\text{SO}_4^{2-}$  is called the sulfide ion.

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### Chapter 18 Nuclear Chemistry

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### Chapter 6: Chemical Composition

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