

Name _____ Date _____ Class _____

CHAPTER 7 REVIEW*Chemical Formulas and Chemical Compounds***SECTION 7-1****SHORT ANSWER** Answer the following questions in the space provided.

1. c In a Stock name such as iron(III) sulfate, the roman numeral tells us _____
- how many atoms of Fe are in one formula unit
 - how many sulfate ions can be attached to the iron atom
 - the charge on each Fe ion
 - the total positive charge of the formula unit
2. c The result of changing a subscript in a correctly written chemical formula is to _____
- change the number of moles represented by the formula
 - change the charges on the other ions in the compound
 - change the formula so that it no longer represents the compound it previously represented
 - have no effect on the formula
3. The explosive TNT has the molecular formula $C_7H_5(NO_2)_3$.
- 4 elements How many elements make up this compound?
 - 6 oxygen atoms How many oxygen atoms are present in one molecule of $C_7H_5(NO_2)_3$?
 - 21 atoms How many atoms in total are present in one molecule of $C_7H_5(NO_2)_3$?
 - 4.2×10^{24} atoms How many atoms are present in a sample of 2×10^{23} molecules of $C_7H_5(NO_2)_3$?
4. How many atoms are present in each of these formula units?
- 11 atoms $Ca(HCO_3)_2$
 - 45 atoms $C_{12}H_{22}O_{11}$
 - 10 atoms $Fe(ClO_2)_3$
 - 9 atoms $Fe(ClO_3)_2$
5. N_2O_5 a. What is the formula for the compound dinitrogen pentoxide?
- carbon(IV) sulfide b. What is the Stock name for the covalent compound CS_2 ?
- H_2SO_3 c. What is the formula for sulfurous acid?
- phosphoric acid d. What is the name for the acid H_3PO_4 ?

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