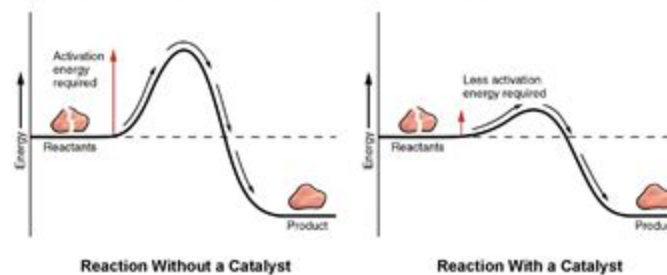


- **Adding catalysts.**

A catalyst is a substance that increases the rate of a chemical reaction without being consumed during the process.

Catalysts function by providing an alternative, lower energy pathway for the formation of the activated complex, thereby lowering the activation energy requirement for the reaction. Greater numbers of particles will now have the kinetic energy needed to react.



Note:

The effects of catalysts can be compounded by coating them onto materials with a high surface area. The high surface area ensures that more reactants are in contact with the catalysts, and therefore, reaction rates are increased. For example, in the production of sulfuric acid, the catalyst V_2O_5 is spread out across large beds so as to maximise the exposed surface area.

IMPORTANT NOTE:

Each of the described changes affects the forward and back reaction rates to the same extent.

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The equilibrium expression for a chemical reaction may be expressed in terms of the concentration of the products and reactants. Only chemical species in the aqueous and gaseous phases are included in the equilibrium expression because the concentrations of liquids and solids does not change. For the chemical reaction: $aA + bB \rightarrow cC + dD$

Chemical Equilibrium in Chemical Reactions

In a chemical reaction, chemical equilibrium is the state in which both reactants and products are present in concentrations which have no further tendency to change with

time, so that there is no observable change in the properties of the system. This state results when the forward reaction proceeds at the same rate as the reverse reaction. The reaction rates of the forward and backward reactions are generally not zero, but equal. Thus, there are no net changes in the ...

Chemical equilibrium - Wikipedia

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Equilibrium | Boundless Chemistry

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