

1<sup>st</sup> Year High School - Chemistry Copy Answer

\*Please write all of your answers on your answer sheet.  
DO NOT WRITE ON THIS TEST!

**Multiple Choice:** Choose the best answer for each of the following questions. Mark your answers on the answer sheet provided.

- What characteristics of aluminum allow it to be made into aluminum foil?
  - It is dense.
  - It is soluble.
  - It is magnetic.
  - It is malleable.
- Which pure substance is composed of more than one element?
  - carbon
  - gold
  - water
  - sodium
- After a chemical reaction occurs in an open test tube, measurements indicate that the mass of the reactants is greater than the mass of the products remaining in the test tube. What **most likely** explains the measurements?
  - A gas was produced.
  - Acid was produced.
  - The acidity of the reactants was high.
  - The temperature of the reactants was low.
- Which factor **best** explains why some people will develop cancer and others will not after being exposed to the same high levels of a harmful chemical?
  - concentration of the chemical
  - potency of the chemical
  - individual susceptibility to the chemical
  - individual exposure to the chemical
- Which is the **best** way to determine if an object is made of pure silver?
  - determine the solubility of the object
  - test to see if the object is chemically reactive
  - determine the density of the object, and compare the known density of pure silver
  - compare the mass of the silver object to the mass of a piece of pure silver
- Which is a naturally occurring chemical?
  - aspirin
  - salt
  - plastic
  - nylon
- Which element is a poor conductor of heat?
  - P
  - K
  - Fe
  - Ag
- How does exposure to cigarette smoke contribute to an asthma attack?
  - by decreasing blood pressure

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# Chemistry Unit 8 Test A Answers

## Chemistry Unit 8 Test A

### Chemistry Unit 8 Test A

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Q. APMQ: A student uses visible spectrophotometry to determine the concentration of  $\text{CoCl}_2$  (aq) in a sample solution. First the student prepares a set of  $\text{CoCl}_2$  (aq) solutions of known concentration. Then the student uses a spectrophotometer to determine the absorbance of each of the standard solutions at a wavelength of 510nm and constructs a standard curve.

### AP Chemistry Unit 8 Electrons and Periodicity PRE/POST ...

Chemistry Unit 8 Test. Study guide. STUDY. PLAY. Dobereiner. grouped elements into 3's based on similar properties (observation of behavior) - called triads. Bunsen. invented the spectroscope; he pioneered the study of elements using their light emitted (1850); he made it easier to identify elements and built a dynasty of science at the University of Heidelberg . Newlands. arranged the ...

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Chemistry Unit 8 Test Questions. Flashcard maker : Lily Taylor. Covalent/molecular compound. containing two or more nonmetals sharing electrons. covalent bond. bond formed when 2 electrons are shared by two nonmetals. molecule. a molecule is the smallest particle in a covalent compound sample. Lewis structural formula . arrangement of electrons around atoms to show bonding. Exothermic. process ...

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SCIENCE 8 CHEMISTRY UNIT TEST REVIEW PACKAGE NAME: \_\_\_\_ The Unit test date is: \_\_\_\_ This package is only a START. Once you have gone through this package, then you should review your notes IN DETAIL, using this package as a guide for the kinds of questions to prepare for, and as a focus for those sections you need to spend more time on. 1. UNDERSTANDING MATTER. Write the meaning of each of ...

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exist throughout nature. T/F. true . The following are properties of compounds \_\_\_\_ . 1: vapor pressure 2: density 3: boiling point 4: all of the above. 4. Solutions are a mixture of 2 or more compounds. T/F. true. A mixture ...

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Q. Two clear colorless solutions were combined. The student observed that the temperature changed from 25 °C to 23 °C when the solutions were combined and that a white substance rapidly formed and settled to the bottom of the container.

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4. If 5.6 g of copper (II) oxide are reacted with 8.6 g of hydrogen according to the following BALANCED reaction, how many grams of copper metal will be produced?  $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$  The limiting reactant is \_\_\_\_, and the excess reactant is \_\_\_\_.

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### Multiple Choice Answers Test ID B 1. D 1. C 2. D 2. C 3. C ...

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