

ChemQuest 29

Balancing Equations

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Introduction Questions

- Write the equation for the reaction of sodium and chlorine (chlorine) to form sodium chloride.
 $2NaCl \rightarrow 2Na + Cl_2$
- Write the equation for the reaction of sodium chloride and sodium chloride to produce sodium chloride and sodium chloride.
 $Ca_3N_2 + 6NaCl \rightarrow 3CaCl_2 + 2Na_3N$

Information: Subscripts and Coefficients

A subscript is a small number that tells you how many atoms are in a compound. For example, in $CaCl_2$, the two is the subscript and it tells us that there are two chlorine ions bonded to one calcium.

A coefficient tells us how many atoms or compounds there are, but is a different way. For example, in the expression "3 NaCl" the three is the coefficient. The three tells us that there are three molecules of water present. In the expression "3 H₂O" there are a total of 6 hydrogen atoms and 3 oxygen atoms.

Critical Thinking Questions

- Verify that in $Al_2(SO_4)_3$ there are 12 oxygen atoms present. $(4 \times 3) \times 4 = 12$
- How many oxygen atoms are in each of the following:
 $3 \text{ atoms } 3 \text{ atoms } 3 \text{ atoms } 20 \text{ atoms}$

Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride from gaseous sodium.
 $Na + Cl_2 \rightarrow NaCl$

Remember that when chlorine is by itself it is always written as Cl_2 . On the reactant side of the reaction (left side) there are a total of two atoms of chlorine, but on the product side there is only one atom of chlorine. Atoms cannot simply disappear. In order for the equation to make sense, we need to balance the equation. This can be done, first, by adding a "2" to the product side.
 $Na + Cl_2 \rightarrow 2NaCl$

Now the equation reads that one atom of the reactant with one molecule of Cl_2 to produce two units of $NaCl$. However, now the Na atoms are not balanced because there is one atom of the reactant side, but two atoms of Na on the product side. This can be fixed by adding another two.
 $2Na + Cl_2 \rightarrow 2NaCl$

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