

**Critical Thinking Questions**

7. Verify that it takes 4140.7 J of energy to heat 12.7 g of water from room temperature (22°C) to the boiling point (100°C). Recall that the specific heat ( $C_p$ ) of water is 4.18 J(g·°C).  
 $\Delta T = 100 - 22 = 78$   
 $q = (m)(C_p)(\Delta T) = (12.7)(4.18)(78) = 4140.7 \text{ J}$
8. Given the fact that the  $\Delta H_{\text{vap}}$  of water is 2260 J/g we know that 2260 J of energy is required to vaporize 1 gram of water. To vaporize 2 grams of water it requires 4520 J. For 3 grams, 6780 J are needed. Verify that it takes 28,702 J of energy to vaporize 12.7 g of water, when the water is already at its boiling point.  
 $(2260 \text{ J/g})(12.7 \text{ g}) = 28,702 \text{ J}$ . note: a helpful equation for phase changes is  $q=(m)(\Delta H)$
9. Verify that it takes 1513.6 J of energy to heat 12.7 g of steam from 100°C to 159°C. Note: the specific heat of steam ( $C_p$ ) is 2.02 J(g·°C).  
 $\Delta T = 159 - 100 = 59$   
 $q = (m)(C_p)(\Delta T) = (12.7)(2.02)(59) = 1513.6 \text{ J}$
10. Using only your answers to questions 7-9, calculate how many Joules of energy it takes to change 12.7 g of water at 22°C to steam at 159°C.  
 Questions 7-9 are each different "steps" for taking 12.7g of water from 22°C to 159°C so you can simply add the answers:  $4140.7 + 28,702 + 1513.6 = 34,356.3 \text{ J}$
11. Calculate how many Joules of energy would be required to change 32.9 g of water at 35°C to steam at 120°C. You will need to break this problem into four steps.  
 a) Find the Joules needed to heat the water to the boiling point.  
 $q = (m)(C_p)(\Delta T) = (32.9)(4.18)(65) = 8938.9 \text{ J}$   
 b) Find the Joules needed to vaporize the water.  
 $q=(m)(\Delta H) = (32.9)(2260) = 74,354 \text{ J}$   
 c) Find the Joules needed to heat the vapor (steam) from the boiling point to 120°C.  
 $q = (m)(C_p)(\Delta T) = (32.9)(2.02)(20) = 1329.2 \text{ J}$   
 d) Add your answers to steps a, b, and c.  
 $8938.9 + 74,354 + 1329.2 = 84,622.1 \text{ J}$
12. How much heat energy would be required to change the temperature of 125g of ice from -32.9°C to liquid water at 75°C?  
 $q = (m)(C_p)(\Delta T) = (125)(2.06)(32.9) = 8471.75 \text{ J}$   
 $q=(m)(\Delta H) = (125)(334) = 41,750 \text{ J}$   
 $q = (m)(C_p)(\Delta T) = (125)(4.18)(75) = 39,187.5 \text{ J}$   
 Total:  $8471.75 + 41,750 + 39,187.5 = 89,409.25 \text{ J}$

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# Chemquest 31 Using Moles With Formulas Answers

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